**DIGI SCHOOL 2020-1-SK01-KA226-SCH-094350 Chemistry – Chemical reactions**



**Classification by outer changes:**

**Synthesis:** simple substances combine to create a more complex one. For example, two reactants turn into one product:

C + O2 → CO2

**Decomposition:** a more complex substance breaks down into simpler parts. For example, one reactant changes into two products:

2NH3 → 3H2 + N2

**Single replacement:** in a reaction an element (according to the galvanic potential) replaces another in a compound:

Na + AgNO3 → Ag + NaNO3

**Double replacement**: in a reaction the cations and anions of two compounds switch places. Neutralisation and precipitation are double replacement reactions, e.g.,   
 Na + AgNO3 → Ag + NaNO3